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ferrous salts in the presence of ferric by this test? c. Acidify 5 cc. of a solution of ferric chloride with hydrochloric acid and allow it to react with metallic aluminum turnings for a few minutes. Decant the liquid from the unused turnings and add to it a few drops of 3 % potassium ferricyanide. (?) What substance must be present? What reaction has the iron undergone? What change has the aluminum undergone? (T. 336.) d. Repeat c but use zinc instead of aluminum. e. Repeat c but use magnesium instead of aluminum. 92. Chlorine and Bromine as Oxidizing Agents (T. 332).--a. Add sodium hydroxide in excess to 5 cc. of a ferrous salt solution. (?) (T. 173.) To 5 cc. more of the same ferrous salt solution add dilute sulfuric acid and then chlorine water until the solution will bleach litmus paper. Then add sodium hydroxide. (If L. 64 was not done by you, add sodium hydroxide to 5 cc. of ferric chloride solution and compare the result with that above.) b. To a solution of mercurous nitrate add chlorine water. What is the precipitate which first forms? (T. 180, L. 68d.) Add more chlorine water until the white precipitate dissolves. Finally boil the solution until it no longer smells of chlorine. Cool the solution and add sodium hydroxide to it.(?) If you did not carry out expt. L. 68d, try the effect of sodium hydroxide on solutions of mercurous and mercuric salts respectively. What salt is present after the action of the chlorine? Formulate the equation bearing in mind that there will be both chloride and nitrate...

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